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Practice Problems
with Step by Step
Answers Orbitals,
Quantum Numbers

\u0026amp; Electron

Configuration -

Multiple Choice

Practice Problems

How to Write the

Electron Configuration

for an Element in

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~~Problem: Electron~~

~~Configuration and~~

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Notation Orbital

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Orbitals, and Electron

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Electronegativity,

Ionization Energy,

Atomic Radius -

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Only the Periodic
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Atomic Radius,

Ionization Energy,

and Electronegativity

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Writing Electron

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Configuration of Ions

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#4 Electron

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electrons || IIT A level

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- Electron

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~~Configuration Test~~

~~With Answers~~

Test Questions and

Answers. 1. What

atom matches this

electron

configuration?

$1s^2 2s^2 2p^6 3s^2$. Neon;

Magnesium;

Aluminum;

Potassium; 2. What

atom matches this

electron

configuration? $1s^2 2s^2$

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$2p^6 3s^2 3p^6 4s^2 3d^{10}$.

Zinc; Copper; Nickel;
Germanium; 3. What
is the electron

configuration for a
Sulfur atom?

$1s^2 2s^2 2p^6 3p^6$;

$1s^2 2s^2 2p^6 3s^2 3p^6$;

$1s^2 2s^2 2p^6 3s^2 3p^4$;

$3p^4$; 4.

~~Electron Configuration~~

~~Practice: Quiz,~~

~~Answers and Basics~~

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Answers . 1. (d) $2n^2$
2. (e) One of five possible values
3. (b) 6 electrons
4. (d) -1, 0, and 1
5. (c) Either set of quantum numbers would express an electron in a 3d orbital
6. (a) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
7. (b) $1s^2 2s^2 2p^6 3s^2 3p^3$
8. (a) $(\uparrow\downarrow) (\uparrow) (\uparrow) (\uparrow) (\uparrow) (\uparrow)$
9.

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~~Electron Configuration~~

~~Test Questions~~

~~ThoughtCo~~

~~Answers~~

The electron configuration of an atom is $1s^2 2s^2 2p^6$. The number of valence electrons in the atom is The number of valence electrons in the atom is answer choices

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~~Configurations |~~

~~Periodic Table Quiz~~

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Answers For Electron

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DRAFT. 10th - 11th
grade. 212 times.

Chemistry. 73%
average accuracy. 8
months ago.

mstruong1. 0. Save.
Edit. Edit. ... Select
answers that DOES
NOT apply to the
quantum-mechanical
model of the atom.
answer choices

~~Unit 4C & 5A Test:~~

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~~Electron~~

~~Configurations and~~

~~Periodic ...~~

~~Answers~~

The electron configuration of an element is shown below. $1s^22s^22p^5$
Name the group this element belongs to in the periodic table and explain your answer.
Based on the electron configuration, expl...

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~~Electron Configuration
Questions and
Answers | Study.com~~

The correct answer to this question is C, $6.165 \times 10^{14} \text{ s}^{-1}$.

This type of question could be found on a chemistry test, as it is related to electron configuration, which is when electrons are distributed throughout an atom. This atom

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configuration.

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configurations.

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~~(practice) | Khan~~

~~Academy~~

Consider the following

electron

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Configurations to

answer the questions
that follow: (i) [Kr] 5s¹

(ii) [Ne] 3s² 3p⁵

(iii) [Ar] 4s² 3d¹⁰ 4p⁴

(iv) [Ne] 3s² 3p⁶ (v)

[Ar] 4s¹ 35) The

electron configuration
of the atom that is

expected to have the
lowest first ionization

energy is _____. A)(i)

B)(ii) C)(iii) D)(iv)

E)(v)

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~~A.P. Chemistry
Practice Test - Ch. 7,
Atomic Structure and~~

...

1.1 Test (mark
scheme) More Exam
Questions on 1.1
Atomic Structure and
1.4 Periodicity ... 1.1
Exercise 2 - ram, rmm
and mass spectra 1.1
Exercise 3 - electronic
configuration 1.1

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Exercise 4 - ionisation energies Answers to 1.1 Exercises. Click here to view some great books which can aid your learning .

~~1.1 Atomic Structure
A Level Chemistry~~

The electron dot structure depends on the number of valence electrons. To answer the question, you

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need to know the electron configuration of the atoms to see which one has 7 unbonded electrons, like chlorine. Fluorine, element number 9, has 2 electrons in the s sublevel (K shell). The L shell is incompletely filled, with 7 electrons.

~~Atomic Structure~~

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~~Chemistry Quiz~~

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the electronic

configuration of

copper atoms is $1s^2$

$2s^2 2p^6 3s^2 3p^6 4s$

$1 3d^{10}$ the electronic

configuration of

copper atoms is $1s^2$

$2s^2 2p^6 3s^2 3p^6 4s$

$2 3d^9$ RE-READ

YOUR NOTES IF

YOU DIDN'T

UNDERSTAND THE

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ANSWERS

Test With

~~ELECTRONIC~~

~~CONFIGURATIONS~~

~~knockhardy.org.uk~~

Answered: Jul 18,
2019. The correct
answer to this

question is C,
 $6.165 \times 10^{14} \text{ s}^{-1}$.

This type of question
could be found on a
chemistry test, as it is
related to electron

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Configuration, which is
when electrons are...

Read More.

Answers

~~Best Electron
Configuration
Questions and
Answers (Q&A ...~~

The electron
configurations of
silicon (14 electrons),
phosphorus (15
electrons), sulfur (16
electrons), chlorine

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(17 electrons), and argon (18 electrons) are analogous in the electron

configurations of their outer shells to their corresponding family members carbon, nitrogen, oxygen, fluorine, and neon, respectively, except that the principal quantum number of the outer shell of the

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heavier elements has increased by one to $n = 3$. Figure  shows the lowest energy, or ground ...

~~2.6: Electron~~

~~Configurations~~

~~Chemistry LibreTexts~~

Dot structures make it easy to count electrons and they show the number of

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Configurations in each electron shell. Arrow and line diagrams show the spin of electrons and show every orbital. Written configurations require minimal space and show the distribution of electrons between subshells.

~~Atomic numbers and electron~~

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assignment and ...

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Gizmo Answers

Keyrar Create the electron configuration of any element by filling electron orbitals.

Determine the relationship between electron configuration and atomic radius.

Discover trends in atomic radii across

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Answers

periods and down families/groups of the periodic table.

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~~Gizmo Answer Key~~

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configurations for the third and fourth

periods. Electron

configurations of the

3d transition metals.

Paramagnetism and

diamagnetism.

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configurations article.
Next lesson. Periodic
table. Test prep ...

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questions (practice) |
Khan Academy~~

38. The ground state
electron configuration
of Fe is _____. A. 2
 $1s^2 2s^2 3s^2 3p^6 3d^6$ B.
 $2 1s^2 2s^2 2p^6 3s^2 3p^6$
 $4s^2 3d^6$ C. $2 1s^2 2s$

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2p⁶ 3s² 3p⁶ 4s² D. 2

1s² 2s² 2p⁶ 3s² 3p⁶

4s² 4d⁶ 39. The

ground state electron configuration of Ga is _____.

A. 2 1s² 2s²

3s² 3p⁶ 3d¹⁰ 4s² 4p¹

B. 6 1s² 2s² 2p⁶ 3s²

3p⁶ 4s² 4d¹⁰ 4p¹ C.

2 1s² 2s² 2p⁶ 3s² 3p⁶

4s² 3d¹⁰ 4p¹

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Electron~~

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Multiple....

A condensed electron configuration is written so that completely filled principal shells are written as the corresponding noble gas, in square brackets. Answer and Explanation: a.

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510bf3f6442f0e4475c
8f9b1adf7

Answers