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Applications half life calculations *Solving Half Life Problems How Nuclear Power Plants Work / Nuclear Energy (Animation)* Nuclear Chemistry Part 2 - Fusion and Fission: Crash Course Chemistry #39 Half Life Decay $N=N_0e^{-kt}$ (Natural Log)

What are Alpha, Beta and Gamma Decay? *Half-Life Calculations: Radioactive Decay Radiation and Radioactive Decay Pearson Chapter 25: Section 2: Nuclear Transformation Lesson 4 - Introduction to Nuclear Chemistry Chapter 21 - Nuclear Chemistry: Part 1 of 9*

Alpha Decay, Beta Decay, Gamma Decay - Electron Capture, Positron Production - Nuclear Chemistry
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25 Nuclear Chemistry Practice

CHAPTER 25 NUCLEAR CHEMISTRY PRACTICE

EXAMPLES 1A (E) A has a mass number of zero and an
“atomic number” of 1. Emission of this electron has
the effect of transforming a neutron into a proton. 94
241 95 241 1 Pu Am +0 1B (E) 58 Ni has a mass
number of 58 and an atomic number of 28. A positron
has a mass

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Problems Answer Key

SECTION 25.1 NUCLEAR RADIATION 1. What happens to the mass number and atomic number of an atom that undergoes beta decay? 2.

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Chapter 25: Nuclear Chemistry 1197 (c) We now determine the number of atoms remaining after 16 days. Because two half- lives elapse in 16 days, the number of atoms has been halved twice, to one-fourth (25%) the original number of atoms.

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